

## Lecture 2

Clausius - Clapeyron equation (all know).

$$\frac{de}{dT} = \frac{L_v}{T \Delta V}, \quad \Delta V = \frac{R_v T}{e} \quad \left[ \text{ideal gas law} \right]$$

$L_v = \text{latent heat vaporisation}$

$$e_{sw} = 610.7 e \left[ \frac{L_v}{R_v} \left( \frac{1}{273.15} - \frac{1}{T} \right) \right] \rightarrow \text{Saturation vapour pressure overlaid}$$

why putting liquid in cloud chamber

comes in to saturation.

$$e_{si} = 610.7 e \left[ \frac{L_s}{R_v} \left( \frac{1}{273.15} - \frac{1}{T} \right) \right] \rightarrow \text{Saturation vapour pressure over ice}$$

$$L_s = \text{latent heat sublimation} = 2.8 \times 10^6 \text{ J kg}^{-1}$$

saturation ratio :-

$$s_e = \frac{e}{e_{sw}}$$

$\rightarrow$  saturation ratio ~~over~~ wrt liquid water

$$s_i = \frac{e}{e_{si}}$$

$\rightarrow$  sat. rat. wrt ice.

Supersaturation :-

$$S_e = s_e - 1$$

$$S_i = s_i - 1$$

$s > 0$  slight required for cloud formation

Example :- You say?

$$W_s = \frac{\epsilon \epsilon_s(T)}{P}$$

difference / which do you take away from which?  
 If I calculate the  $V W_s$  at two levels that will give adiabatic LWMR

$$W_{s1} = \epsilon \times 610.7 e \left[ \frac{2.8 \times 10^6}{461} \left( \frac{1}{273.15} - \frac{1}{280.15} \right) \right] \quad \left[ \epsilon = \frac{R_g}{R_v} = 0.622 \right]$$

$$= \frac{900 \times 10^2}{6.93 \times 10^{-3}} \times 7.36 \times 10^{-3} \text{ kg kg}^{-1}$$

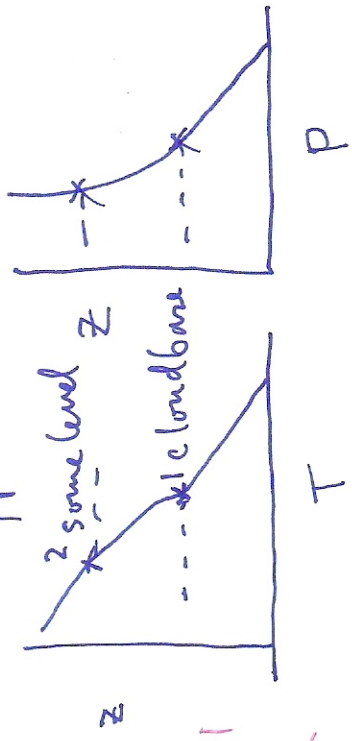
$$W_{s2} = \epsilon \times 610.7 e \left[ \frac{2.8 \times 10^6}{461} \left( \frac{1}{273.15} - \frac{1}{278.65} \right) \right]$$

$$= \frac{6.17 \times 10^{-3}}{6.4 \times 10^{-3}} \times 6.7 \times 10^{-1} \text{ kg kg}^{-1}$$

$$= \frac{0.27 \times 10^{-3} \text{ kg/kg}}{0.8 \times 10^{-3}}$$

$$\therefore W_{s1} - W_{s2} \approx 0.8 \times 10^{-3} \quad \underline{0.23 \text{ kg kg}^{-1}}$$

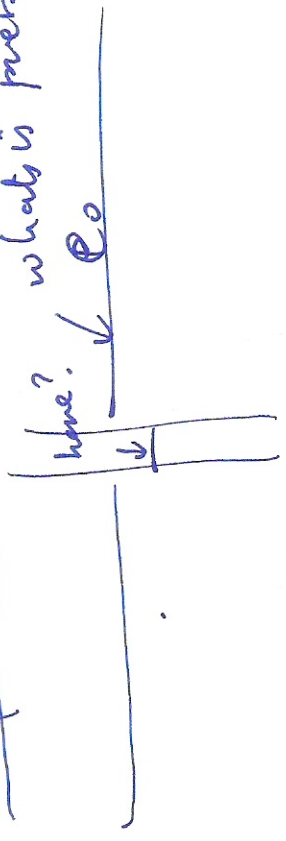
Approx.



KAM  
 $ALWMR = W_{s,1} - W_{s,2}$

Also web page

Recap on flat surfaces Q1.



Why is water in tube below water level?

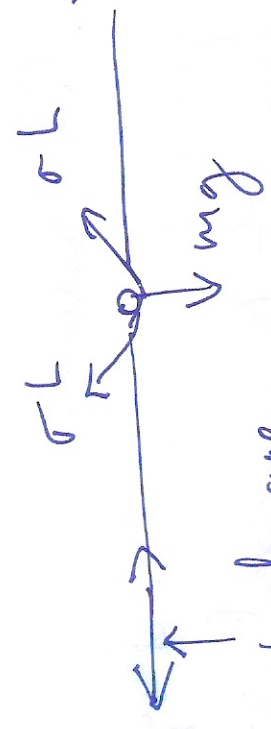
— take answers

Q2.

Why does the pin stay a float?  
→ displaces water that is heavier than it?

Archimedes - No!

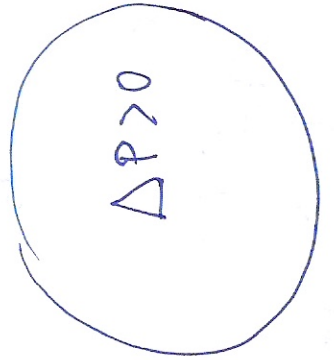
Steel is denser than water and the pin is on top of water.



molecules are attracted to each other.

Surface Tension

Q3

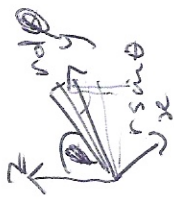


— Balloon  
Any ideas on how I could calculate ΔP from shape? What would I need?

Derive it.

Consider a capillary tube - why?

- turns out to be relevant.



$$dA = r^2 \sin\theta d\theta d\phi$$

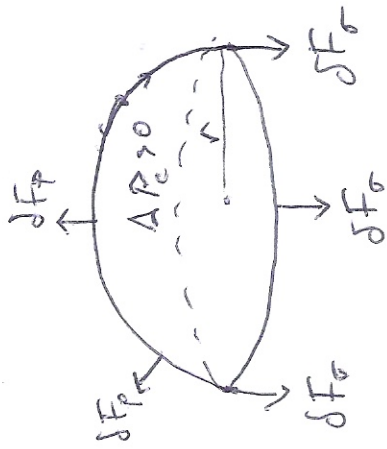
$$\delta F_p = \Delta P_c r^2 \sin\theta d\theta d\phi$$

but we want component in z dir<sup>n</sup>

$$\delta F_{pz} = \Delta P_c r^2 \sin\theta d\theta d\phi \times \cos\theta$$

$$F_p = \int_{\phi=0}^{2\pi} \int_{\theta=0}^{\pi/2} \Delta P_c r^2 \sin\theta \cos\theta d\theta d\phi$$

$$= \frac{\Delta P_c \pi r^2}{2}$$



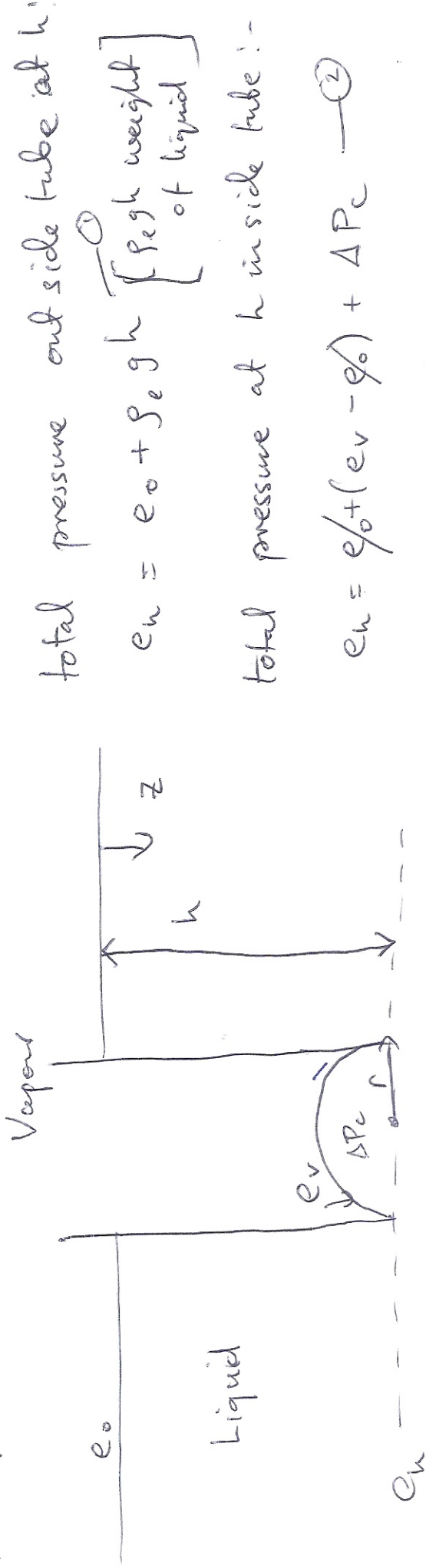
$$F_\sigma = 2\pi r \sigma$$

$$\Delta P_c r = 2\pi r \sigma$$

$$\Delta P_c = \frac{2\sigma}{r}$$

[ Pressure inside drop ]  
[ Laplace pressure. ]

Capillary tubes - why? it turns out to be relevant.



total pressure out side tube at  $h$ :

$$e_h = e_0 + \rho_l g h \quad \left[ \rho_l g \text{ weight of liquid} \right] \quad \textcircled{1}$$

total pressure at  $h$  inside tube:-

$$e_h = e_0 + (e_v - e_0) + \Delta P_c \quad \textcircled{2}$$

hydrostatic relation for vapor:  $\left[ \frac{de}{dz} = \frac{\rho}{RvT} g \right]$

$$e_v = e_0 \exp\left(\frac{g h}{RvT}\right) \quad \textcircled{3}$$

~~$$e_0 + \rho_l g h = e_0 \exp\left(\frac{g h}{RvT}\right) + \Delta P_c$$~~

Combine  $\textcircled{1}$  +  $\textcircled{2}$  &  $e_0 \approx e_v$

~~$$e_0 + \rho_l g h = e_0 \exp\left(\frac{g h}{RvT}\right) + \Delta P_c$$~~

$$\Rightarrow h \approx \frac{\Delta P_c}{\rho_l g} \quad \textcircled{4}$$

$\textcircled{4}$  in  $\textcircled{3}$ :

$$e_v = e_0 \exp\left(\frac{\Delta P_c}{\rho_l RvT}\right)$$

Sub  $\Delta P_c$ :

$$e_v = e_0 \exp\left(\frac{2\sigma}{\rho_l RvT r}\right)$$

- also William Thompson.

Kelvin's equation - weird?

Vapor pressure required to support small drops depends on size?

- Kelvin's equation

## Key points.

- Saturation vapor pressure over liquid (water)
- Saturation vapor pressure over ice
- Saturation ratio & super-saturation of water vapor
- Why size is important for the growth of aerosol particles into cloud drops - Kelvin's equation.